The Chemistry of Life Section 2–1 The Nature of Matter (pages 35–39) This section identifies the three particles that make up atoms. It also explains how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds. Atoms (page 35) 1. The basic unit of matter is called a(an) 2. Describe the nucleus of an atom. 3. Complete the table about subatomic particles. SUBATOMIC PARTICLES Particle Charge Location in Atom Positive Neutral Negative 4. Why are atoms neutral despite having charged particles? Elements and Isotopes (page 36) 5. What is a chemical element? 6. What does an element's atomic number represent? 7. Atoms of the same element that differ in the number of neutrons they contain are known as 8. How are isotopes identified? 9. Why do all isotopes of an element have the same chemical	Vame	Class	Date
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Vame	Class	Date
Chapter 2, The	Chemistry of Life (continued)	
compound?	he formula for table salt indicate about that	
<u></u>	•	
Chemical Bo	nds (pages 38–39)	
	atoms in compounds together?	
	ne table about the main types of chemical bonds.	
is. Complete tr	CHEMICAL BONDS	
Туре	Formed when	
Covalent bond		
lonic bond		
14. What is an i	on?	
	ving sentence true or false? An atom that loses	
electrons ha	s a negative charge	
	re that results when atoms are joined together by	
covalent bo		
	nds is called a(an)	
17. Circle the le	nds is called a(an) etter of each sentence that is true about covalent bonds.	
	·	
a. When at	etter of each sentence that is true about covalent bonds. coms share two electrons, it is called a double bond. er molecule, each hydrogen atom forms a single	
a. When atob. In a watecovalent	etter of each sentence that is true about covalent bonds. coms share two electrons, it is called a double bond. er molecule, each hydrogen atom forms a single	
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a. When atob. In a watecovalentc. Atoms cad. In a cova	etter of each sentence that is true about covalent bonds. coms share two electrons, it is called a double bond. er molecule, each hydrogen atom forms a single bond. bond. an share six electrons and form a triple bond.	

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